Unit 6 Nomenclature

1. Determine if the follow metals have only one charge or more than one charge
   a. Sr
   b. Ti
   c. Y
   d. Ni
2. Determine the charge for each of the following mono-atomic anions
   a. sulfide
   b. iodide
   c. phosphide
   d. oxide
3. Name the following
   a. $CaBr_2$
   b. $FeCl_3$
   c. Cu_2S
   d. RbI
   e. Au_2O_3
   f. Ba_3N_2
4. Write formulas for the following
   a. cobalt (III) chloride
   b. strontium iodide
   c. silver phosphide
   d. palladium (IV) sulfide
   e. osmium (III) bromide
   f. zirconium oxide
5. Name the following
   a. Na_2SO_4
   b. $Fe(NO_3)_2$
   c. $Cu(ClO_4)_2$
   d. $V_3(PO_4)_4$
   e. $Mn_2(CO_3)_5$
   f. K_2SO_3
6. Write formulas for the following
   a. zinc nitrite
   b. barium chromate
   c. nickel (III) iodate
   d. tungsten (V) hydroxide
   e. lead (IV) carbonate
   f. calcium hypobromite

7. Name the following acids
    a. $HBrO_3$
    b. H_2SO_4
    c. HNO_3
    d. HCN
    e. H_3PO_3
    f. H_2CO_3
8. Write formulas for the following
    a. nitrous acid
    b. hydroiodic acid
    c. perchloric acid
    d. hypobromous acid
    e. sulfurous acid
    f. chloric acid
9. Name the following
    a. Cl_2O
    b. OF_2
    c. CCl_4
    d. P_2S_5
    e. NO_3
    f. S_3Br_8
10. Name or write formulas for the following
    a. $Na_2CO_3 \cdot 4\ H_2O$
    b. potassium sulfate pentahydrate
11. Fill in the following chart for practice. Write each formula and name each compound.

Cl^{-1} NO_3^{-1} O^{-2} SO_3^{-2} PO_4^{-3} CO_3^{-2} N^{-3}

Ba^{+2}

Al^{+3}

K^{+1}

Fe^{+2}

Cu^{+1}

Ta^{+5}